

The rules for working out oxidation states in compounds and complex ions

Element	Oxidation state	Exception(s)	What to do
Oxygen	-2	When combined with F or in the peroxide ion (O₂) .	The oxidation state must balance accordingly even if this is not the usual rule. The other species has preference.
Chlorine	-1	When combined with O or F (more electronegative)	
Bromine	-1	When combined with O, F or Cl (more electronegative)	
Iodine	-1	When combined with O, F, Cl or Br (more electronegative)	
Hydrogen	+1	When in a metal hydride (e.g. NaH)	
Fluorine	-1		
Group 1	+1		
Group 2	+2		
Aluminium	+3		

